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My background is in elementary particle physics and accelerator physics, but nowadays I am better known as a textbook author and for various editorial roles at the American Journal of Physics. I have worked with undergraduate students on a variety of projects in quantum mechanics, thermal physics, relativity, and other areas of theoretical physics.

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$V(2) = m kT P(3)$  where  $m$  is the average mass of an air molecule,  $N$  is the number of molecules in volume  $V$  and  $T$  is the temperature. We therefore get a differential equation for the pressure:  $dP/dz = -\rho g(4) = -\rho mg/kT P(5)$  This is called the barometric equation. In a realistic model, both  $\rho$  and  $P$  are functions of  $z$ .

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